

ever, glycine-containing substrates are split under these conditions since glycine contains no side chain in such a spatial arrangement that it could prevent an approach of the enzyme.

It is a pleasure to acknowledge many valuable suggestions which Dr. Max Bergmann has offered in a discussion of this problem.

DEPARTMENT OF BIOCHEMISTRY
SCHOOL OF MEDICINE
DUKE UNIVERSITY
DURHAM, NORTH CAROLINA

RECEIVED JULY 23, 1943

NEW COMPOUND

4-Phenylphenyl Butyrate

This compound was prepared in 81% yield by the inter-

action of 4-phenylphenol and butyryl chloride¹ in pyridine solution with 1,4-dioxane as diluent.² The crude product was dissolved in benzene, and the resulting solution was washed with dilute hydrochloric acid and sodium hydroxide solution and decolorized with Norite. After removal of the benzene on the steam-bath, the product was crystallized four times from 30–60° ligroin; colorless platelets resulted; m. p. 59–60.3°.

Anal. Calcd. for C₁₈H₁₈O₂: C, 80.0; H, 6.67. Found: C, 79.83; H, 6.85.

(1) Gilman and Blatt, "Organic Syntheses," John Wiley and Sons, Inc., New York, N. Y., Coll. Vol. I, 2d ed., 1941, p. 147.

(2) Hazlet, Hensley and Jass, *THIS JOURNAL*, **64**, 2449 (1942).

DEPARTMENT OF CHEMISTRY
STATE COLLEGE OF WASHINGTON
PULLMAN, WASHINGTON

STEWART E. HAZLET
LEE C. HENSLEY

RECEIVED JULY 20, 1943

COMMUNICATION TO THE EDITOR

THERMAL PROPERTIES OF ISOPENTANE

Sir:

Anyone not completely familiar with the field, on reading the paper by Guthrie and Huffman on page 1143 of the June, 1943, issue of *THIS JOURNAL* might receive a wrong impression. We have carefully reviewed our work published on isopentane in *THIS JOURNAL* and find no reason to doubt the data there reported. The facts are that two independent groups of workers using different calorimeters have observed an anomaly in the thermal behavior of isopentane. A third group of workers using a different calorimeter

have failed to observe this phenomenon. The second independent series of measurements was made in our laboratory but in a different calorimeter (Gold calorimeter C). The work was done by M. L. Sagenkahn and H. F. Zuhr. The third independent series of measurements is that of Guthrie and Huffman. After the present emergency we shall repeat again the work on isopentane using the Huffman type calorimeter.

SCHOOL OF CHEMISTRY AND PHYSICS
THE PENNSYLVANIA STATE COLLEGE
STATE COLLEGE, PA.

J. G. ASTON

RECEIVED AUGUST 6, 1943

NEW BOOKS

Organic Syntheses. Volume 23. LEE IRVIN SMITH, Editor-in-Chief, HOMER ADKINS, C. F. H. ALLEN, W. E. BACHMANN, NATHAN L. DRAKE, C. S. HAMILTON, R. L. SHRINER, H. R. SNYDER AND A. H. BLATT, Secretary to the Board. John Wiley and Sons, Inc., 440 Fourth Avenue, New York, N. Y., 1943. 124 pp. 15 × 23.5 cm. Price, \$1.75.

In this last volume of this important series of *Organic Syntheses* are recorded specific directions for preparing thirty-nine different organic compounds. These embrace representatives of the aliphatic, aromatic and heterocyclic series; and fifty-four contributors other than

members of the publishing Board have taken part in the development and construction of the experimental technique described. The techniques proposed have been checked in each experiment by two independent workers. A short literature review accompanies each preparation and in many cases useful notes are inserted which serve to guide the experimenter in applying the experimental procedure recommended. All the procedures are clearly written, and the book should find a most useful service in every laboratory where organic synthesis is being applied and practiced.

TREAT B. JOHNSON